

## $^{82}\text{Se}$

Aston discovered  $^{82}\text{Se}$  in 1922 as reported in “The Isotopes of Selenium and some other Elements” ([1922As02](#)). Selenium was vaporized in a discharge tube to obtain suitable spectra with the Cavendish mass spectrometer. “The interpretation of these is quite simple and definite, so that the results may be stated with every confidence. Selenium consists of six isotopes, giving lines at 74(f), 76(c), 77(e), 78(b), 80(a), 82(d). The line at 74 is extremely faint. The intensities of the lines are in the order indicated by the letters, and agree well enough with the chemical atomic weight 79.2.”

Adapted from reference ([2012Gr02](#))

[1922As02](#) F. W. Aston, *Nature* **110**, 664 (1922).

[2012Gr02](#) J. L. Gross, J. Claes, J. Kathawa, and M. Thoennessen, *At. Data Nucl. Data Tables* **98**, 75 (2012).

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